

Atomic Structure Chapter 4 Answers

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(i) Atomic number = number of protons + number of electrons (ii) Mass number = number of protons + number of neutrons (iii) Atomic mass = number of protons = number of neutrons (iv) Atomic number = number of protons = number of electrons (a) (i) and (ii) (b) (i) and (iii) (c) (ii) and (iii) (d) (ii) and (iv) Answer: Answer: (d) (ii) and (iv)

MCQ Questions for Class 9 Science Chapter 4 Structure of ...
Section 4.2Structure of the Nuclear Atom. One change to Dalton's atomic theory is that . atoms are divisible. into subatomic particles. Electrons, protons, and neutrons are examples of these fundamental particles. There are many other types of particles, but we will study these three

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The atom of an element is made up of 4 protons, 5 neutrons and 4 electrons. What is its atomic number and mass number? Answer: Protons = 4, neutrons = 5, electrons = 4 Atomic number = 4, Mass number = 4 + 5 = 9 Question 16. (a) What are the two main parts of which an atom is made of? (b) Where is the nucleus of an atom situated? (c) What are ...

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CHAPTER 4 TEST: Atoms, Atomic Theory and Atomic Structure Matching. A. Bohr B. Democritus C. Rutherford D. Dalton E. Thomson F. Schrodinger ____ 1. Greek thinker, called nature's basic particle an atom, based on the Greek word "atomos" which means "indivisible". Did not have evidence that atoms existed. ____ 2.

CHAPTER 4 TEST: Atoms, Atomic Theory and Atomic Structure
Unit 3 Review - Chapters 4 (Atomic Structure) & 6 (Periodic Table) Part 1: Answer the following questions. 1. a. Which scientist created the first modern atomic theory? b. What was his theory? c. What was his model of the atom called?_ d. What was proven wrong and why? 2. a. Which scientist did the gold foil experiment? b.

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